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Adding the number of protons (12) and the number of neutrons (11) together, reveals the mass of the atom to be 23. Many excited state configurations are possible. A few examples include: 2-7-4, 2-5-6, 1-8-4. They contain different numbers of neutrons.

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CH 41. Butanol or n-butanol 42.  
[1] Allow 1 credit. Examples of a  
1-credit response: 12. 3 43. 36  
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A solution was made by dissolving  
12.5 g of  $\text{Na}_3\text{PO}_4$  in 125.00 mL of  
water. The volume of the  
resulting solution was 128 mL.  
Calculate the molarity and  
molality of the solution.

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100 Both isotopes contain 6 protons but different numbers of neutrons. Or: An isotope with a mass of 12.00 contains 6 protons and 6 neutrons, while the isotope with a mass of 13.00 contains 6 protons and 7 neutrons.

CHEMISTRY ANSWER KEY 27 Na

Mg Al  $1s^2 2s^2 2p^6 3s^2 3p^1$  Si

$1s^2 2s^2 2p^6 3s^2 3p^2$  P S

$1s^2 2s^2 2p^6 3s^2 3p^4$

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Na Mg 1s<sup>2</sup>2s<sup>2</sup>2p<sup>6</sup>3s<sup>2</sup>

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12 + H<sub>2</sub>O → H<sub>2</sub>O + + 1- The vertical columns of the table should have elements with similar chemical properties. If the elements were arranged by atomic mass, several pairs of elements, such as Ar and K, would be out of position in regards to properties. Both groups follow the octet rule. Group I members achieve an octet by losing an electron,

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